



Determine the conjugate acid for each:



Determine the conjugate base for each:



$$\text{pH} = -\log[\text{H}^+]$$

$$[\text{H}^+] = 10^{-\text{pH}}$$

Calculate the pH for each $[\text{H}^+]$ given below:

$$[\text{H}^+] = 0.00025$$

$$[\text{H}^+] = 1.0 \times 10^{-7}$$

$$[\text{H}^+] = 2.3 \times 10^{-3}$$

$$[\text{H}^+] = 7.6 \times 10^{-12}$$

Determine the $[\text{H}^+]$ from the pH:

$$\text{pH} = 7.4$$

$$\text{pH} = 2.8$$

$$\text{pH} = 12.7$$

$$\text{pH} = 4.5$$