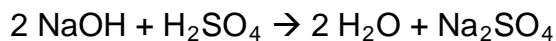


## **Stoichiometry Practice Worksheet**

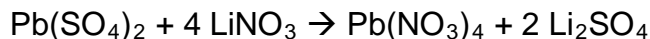
*Solve the following stoichiometry grams-grams problems:*

- 1) Using the following equation:



How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

- 2) Using the following equation:



How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

## **Summer Review Sheet #4**

### *Balancing Equations and Simple Stoichiometry*

*Answers are provided on the second sheet. Please try to do the worksheet without referring to them, because you'll be expected to know this stuff the first day of school!*

Balance the following equations:

- 1)    \_\_\_  $\text{N}_2$  + \_\_\_  $\text{F}_2$   $\rightarrow$  \_\_\_  $\text{NF}_3$
- 2)    \_\_\_  $\text{C}_6\text{H}_{10}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$  + \_\_\_  $\text{H}_2\text{O}$
- 3)    \_\_\_  $\text{HBr}$  + \_\_\_  $\text{KHCO}_3$   $\rightarrow$  \_\_\_  $\text{H}_2\text{O}$  + \_\_\_  $\text{KBr}$  + \_\_\_  $\text{CO}_2$
- 4)    \_\_\_  $\text{GaBr}_3$  + \_\_\_  $\text{Na}_2\text{SO}_3$   $\rightarrow$  \_\_\_  $\text{Ga}_2(\text{SO}_3)_3$  + \_\_\_  $\text{NaBr}$
- 5)    \_\_\_  $\text{SnO}$  + \_\_\_  $\text{NF}_3$   $\rightarrow$  \_\_\_  $\text{SnF}_2$  + \_\_\_  $\text{N}_2\text{O}_3$

Using the equation from problem 2 above, answer the following questions:

- 6)    If I do this reaction with 35 grams of  $\text{C}_6\text{H}_{10}$  and 45 grams of oxygen, how many grams of carbon dioxide will be formed?
  
  
  
  
  
  
  
  
  
  
- 7)    What is the limiting reagent for problem 6? \_\_\_\_\_
- 8)    How much of the excess reagent is left over after the reaction from problem 6 is finished?
  
  
  
  
  
  
  
  
  
  
- 9)    If 35 grams of carbon dioxide are actually formed from the reaction in problem 6, what is the percent yield of this reaction?