

Name _____ **Types of Reactions**

Choose the correct symbol for the type of reaction. Place that answer in the blank at the beginning of each equation and then balance each equation correctly.

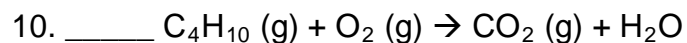
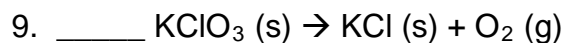
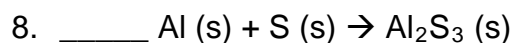
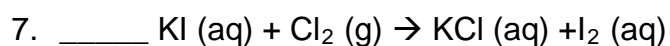
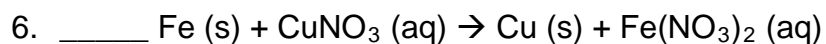
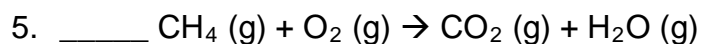
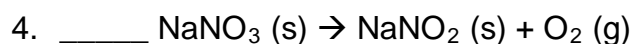
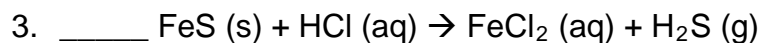
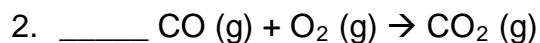
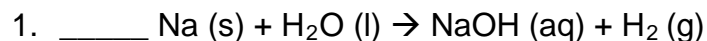
S = synthesis

D = decomposition

C = combustion

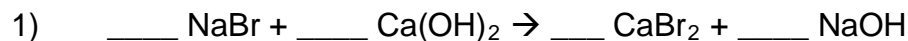
SD = single displacement

DD = double displacement

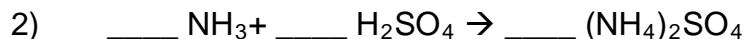


Six Types of Chemical Reaction Worksheet

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:



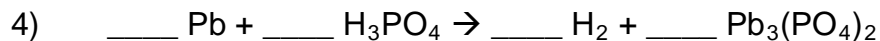
Type of reaction: _____



Type of reaction: _____



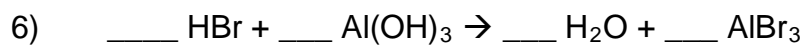
Type of reaction: _____



Type of reaction: _____



Type of reaction: _____



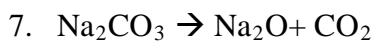
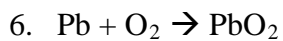
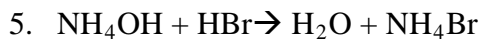
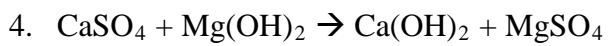
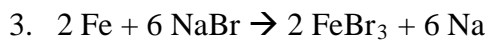
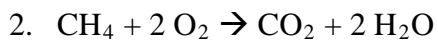
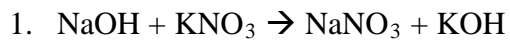
Type of reaction: _____

7) What's the main difference between a double displacement reaction and an acid-base reaction?

8) Combustion reactions always result in the formation of water. What other types of chemical reaction may result in the formation of water? Write examples of these reactions on the opposite side of this paper.

Name: _____

List the type of reaction indicated by the following equations:



Name: _____

Identify the following chemical equations by type.	
1. $\text{C}_4\text{H}_8 + 6\text{O}_2 \rightarrow 4\text{CO}_2 + 4\text{H}_2\text{O}$	2. $\text{HCl} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCl}$
3. $2\text{KNO}_{3(\text{s})} \rightarrow 2\text{KNO}_{2(\text{s})} + \text{O}_{2(\text{g})}$	4. $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{NaNO}_3 + \text{AgCl}$
5. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$	6. $2\text{Ag} + \text{S} \rightarrow \text{Ag}_2\text{S}$
7. $\text{MgCO}_{3(\text{s})} \rightarrow \text{MgO}_{(\text{s})} + \text{CO}_{2(\text{g})}$	8. $\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$