

Explain the following in your own words. (1 point each)

1. What does a **driving force** do?

2. What is an **acid**?

3. What is a **base**?

When electrons are transferred in a reaction, what 2 processes are occurring at the same time?

4. \_\_\_\_\_

5. \_\_\_\_\_

6. Gain of electrons is known as \_\_\_\_\_.

7. Loss of electrons is known as \_\_\_\_\_.

8. When a reaction forms a gas what other **two** driving forces might also be helping to cause the reaction. (half point each)

Name the four driving forces for reactions that we discussed in Chapter 7. (1 point each)

9.

10.

11.

12.

Answer the following with a word or phrase. (1 point each)

13. If a reaction is termed a “REDOX” reaction, then the driving force is the

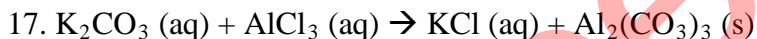
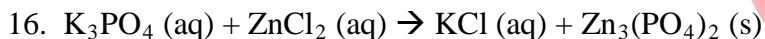
\_\_\_\_\_.

14. An “ACID-BASE” reaction results in the formation of \_\_\_\_\_.

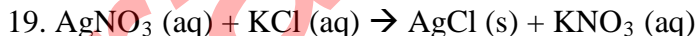
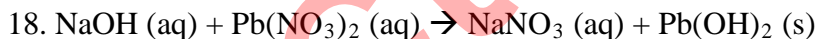
15. A net ionic equation shows only those chemical species that are

\_\_\_\_\_.

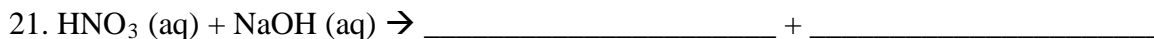
For the following precipitation reaction write the COMPLETE IONIC EQUATION. (4 points)



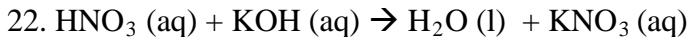
For the following precipitation reaction, write the NET IONIC EQUATION. (3 points)



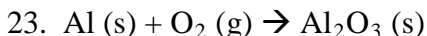
Predict the products for the following reactions. Circle the product that is the salt. (2 points each)



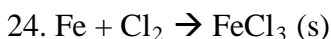
For the following acid-base reaction, write the NET IONIC EQUATION. (3 points)



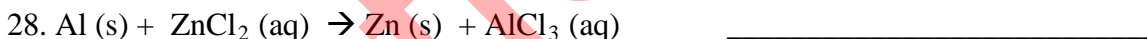
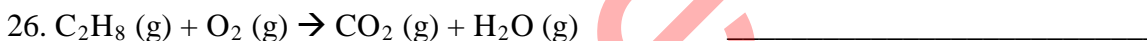
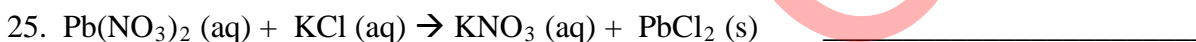
Circle the substance being oxidized in the following unbalanced equation.



Circle the substance being reduced in the following unbalanced equation.



For the following unbalanced reactions identify the type of reaction based on what is happening. Each choice is used only once. Your possible choices are: COMBUSTION, DECOMPOSITION, DOUBLE DISPLACEMENT, SINGLE DISPLACEMENT, SYNTHESIS.



For each of the following reactions identify **two** driving forces for the reaction. (2 pts each)

