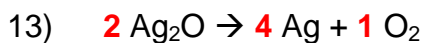
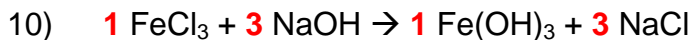
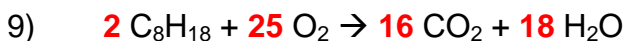
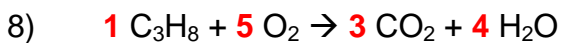
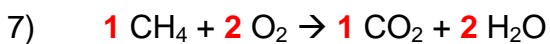
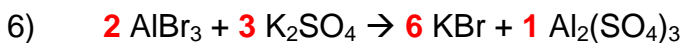
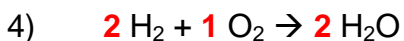
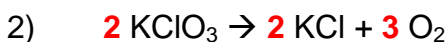
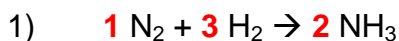


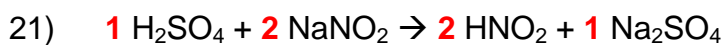
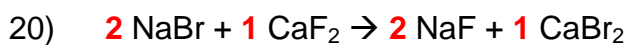
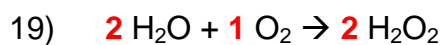
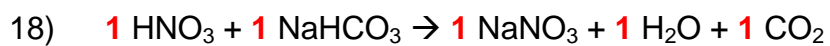
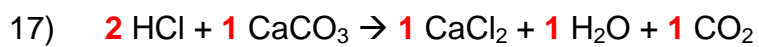
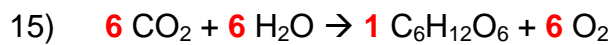
Answers

- 1) $\underline{2} \text{C}_6\text{H}_6 + \underline{15} \text{O}_2 \rightarrow \underline{6} \text{H}_2\text{O} + \underline{12} \text{CO}_2$
- 2) $\underline{4} \text{NaI} + \underline{1} \text{Pb}(\text{SO}_4)_2 \rightarrow \underline{1} \text{PbI}_4 + \underline{2} \text{Na}_2\text{SO}_4$
- 3) $\underline{2} \text{NH}_3 + \underline{2} \text{O}_2 \rightarrow \underline{1} \text{N}_2\text{O} + \underline{3} \text{H}_2\text{O}$
- 4) $\underline{2} \text{Fe}(\text{OH})_3 \rightarrow \underline{1} \text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2\text{O}$
- 5) $\underline{2} \text{HNO}_3 + \underline{1} \text{Mg}(\text{OH})_2 \rightarrow \underline{2} \text{H}_2\text{O} + \underline{1} \text{Mg}(\text{NO}_3)_2$
- 6) $\underline{1} \text{H}_3\text{PO}_4 + \underline{3} \text{NaBr} \rightarrow \underline{3} \text{HBr} + \underline{1} \text{Na}_3\text{PO}_4$
- 7) $\underline{3} \text{C} + \underline{4} \text{H}_2 \rightarrow \underline{1} \text{C}_3\text{H}_8$
- 8) $\underline{2} \text{CaO} + \underline{1} \text{MnI}_4 \rightarrow \underline{1} \text{MnO}_2 + \underline{2} \text{CaI}_2$
- 9) $\underline{1} \text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2\text{O} \rightarrow \underline{2} \text{Fe}(\text{OH})_3$
- 10) $\underline{1} \text{C}_2\text{H}_2 + \underline{2} \text{H}_2 \rightarrow \underline{1} \text{C}_2\text{H}_6$
- 11) $\underline{2} \text{VF}_5 + \underline{10} \text{HI} \rightarrow \underline{1} \text{V}_2\text{I}_{10} + \underline{10} \text{HF}$
- 12) $\underline{1} \text{OsO}_4 + \underline{2} \text{PtCl}_4 \rightarrow \underline{2} \text{PtO}_2 + \underline{1} \text{OsCl}_8$
- 13) $\underline{1} \text{CF}_4 + \underline{2} \text{Br}_2 \rightarrow \underline{1} \text{CBr}_4 + \underline{2} \text{F}_2$
- 14) $\underline{2} \text{Hg}_2\text{I}_2 + \underline{1} \text{O}_2 \rightarrow \underline{2} \text{Hg}_2\text{O} + \underline{2} \text{I}_2$
- 15) $\underline{1} \text{Y}(\text{NO}_3)_3 + \underline{1} \text{GaPO}_4 \rightarrow \underline{1} \text{YPO}_4 + \underline{1} \text{Ga}(\text{NO}_3)_3$

Balancing Chemical Equations – Answers

Balance the equations below:





Balancing Worksheet #1

Answers

1. 2, 1 ----> 2
2. 1, 12 ----> 8
3. 2 ----> 2, 1
4. 1, 2 ----> 1, 1
5. 2, 2 ----> 2, 1
6. 1, 8 ----> 10, 16
7. 4, 11 ----> 8, 6
8. 4, 3 ----> 2
9. 2, 15 ----> 14, 6
10. 4, 11 ----> 2, 8
11. 1, 3 ----> 2, 3
12. 2, 1 ----> 2
13. 2, 5 ----> 4, 2
14. 2 ----> 2, 1
15. 1, 11 ----> 7, 8
16. 1, 4 ----> 1, 2
17. 2 ----> 2, 3
18. 4 ----> 3, 1
19. 1, 6 ----> 4
20. 4, 3 ----> 1
21. 1, 5 ----> 3, 4
22. 1, 3 ----> 2, 3
23. 1, 4 ----> 5, 1
24. 8, 8 ----> 1, 16
25. 3, 4 ----> 1, 4
26. 1, 3 ----> 2
27. 2, 1 ----> 2
28. 6, 6 ----> 1, 6
29. 1, 4 ----> 1, 4
30. 2 ----> 1, 1
31. 1, 2 ----> 1, 1
32. 2, 3 ----> 1, 6
33. 1, 6 ----> 3, 2
34. 1, 8 ----> 1, 4, 4
35. 2, 3 ----> 1, 3
36. 1, 2 ----> 2, 1, 1
37. 1, 5 ----> 2
38. 1, 1 ----> 2
39. 4, 3 ----> 2
40. 2, 2 ----> 4, 1
41. 1, 1 ----> 1, 1
42. 2 ----> 1, 3
43. 1, 3 ----> 2, 3
44. 1, 3 ----> 1, 3
45. 2, 6 ----> 1, 6
46. 1, 1 ----> 2
47. 2, 6 ----> 2, 3
48. 6 ----> 1, 7
49. 3, 1 ----> 1
50. 2, 1 ----> 1, 1, 1

Balance the following equations by placing correct whole number coefficients in the blanks.

